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Complete the following using chapter 5.2 and 5.3

**In order to earn full credit you must write in complete sentences AND state the page number where you found the answer!**

Chapter 5.2

1. Describe the characteristics of salts.
2. Describe why an ionic bond is formed between cations and anions.
3. Do ionic compounds have charges? Why or why not?
4. Explain why ionic bonds are so strong.
5. Salts are both “hard and brittle”. What do each of these terms mean? Why do salts have these properties?
6. What causes the crystal shape that is found in most salts?
7. What distinctive properties do ionic compounds have? (note: you should have three )
8. What is the difference between an ionic compound and a molecule?
9. What two requirements are there for a substance to conduct electricity?
10. What two things affect how the crystal lattice of a salt is formed?
11. Why can salts conduct electricity when they are melted or dissolved but not when they are a solid?
12. You are given an unknown substance. What procedures in the lab could you perform to investigate if it is an ionic compound?

Chapter 5.3

1. Where does a simple cation get its name?
2. How is a simple anion named?
3. How do you indicate the charge of the ion if an element forms two or more ions?
4. What two things make up the name of a binary ionic compound?
5. What is the name of a binary ionic compound made from Na and O?
6. What is the name of a binary ionic compound made from Mg and P?