Your task: Each of the following 20 statements apply to either exothermic, endothermic or both reactions. Cut out the statements and sort them into three piles: **Exothermic vs Endothermic vs Both.** Once you have determined that your statements are in the correct pile, glue them onto the graphic organizer provided.

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| --- | --- | --- | --- |
| Energy is taken in from the surroundings. | Energy is given out to the surroundings. | Products have more energy than reactants. | Products have less energy than reactants. |
| Surrounding loses energy | Can be identified by a change in temperature. | Chemical change. | System loses energy. |
| Surroundings gain energy. | System gains energy. | Surroundings usually feel cool. | Surroundings usually feel warm. |
| Involves heat. | Occurs as chemical bonds break and reform. | No energy is created or destroyed overall. | Melting ice cream. |
| Boiling water. | Instant ice pack. | Instant hand warmers. | Energy change can be measured by temperature. |

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